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 FOR PRODUCTION THEREOF)

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VERIFICATION OF TRANSLATION

Sir:

I, Isamu Ogoshi, having been warned that willful false statements and the like are punishable by fine or imprisonment or both, under section 1001 of Title 18 of the United States Code, and may jeopardize the validity of the above-captioned application and any patent issuing thereon, declare:

- (1) I am a patent attorney authorized to practice law in Japan and am engaged in the practice of law with OGOSHI International Patent Office at Toranomon 9 Mori Bldg. 3F, 2-2, Atago 1-Chome, Minato-ku, Tokyo 105-0002, Japan.
- (2) I am fluent in the Japanese and English Languages.

(3) I have reviewed the attached translation, and certify that it is an accurate English translation of the Japanese language international application of Yuichiro Shindo and Kouichi Takemoto filed on August 12, 2003 and given International Application No. PCT/JP2003/010251.

(4) All of the statements made herein of my own knowledge are true and all statements made herein on information and belief are believed to be true.

January 21, 2005
Date

Isamu Ogoshi
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High Purity Copper Sulfate and Method for Production Thereof

5

TECHNICAL FIELD

The present invention relates to a manufacturing method of high purity copper sulfate which includes the steps of dissolving commercially available 10 copper sulfate crystals (purity, for instance, is 95 to 99.9wt%) in purified water, performing thermal concentration thereto, and filtering out the crystals precipitated initially to eliminate the impurities, and to the high purity copper sulfate obtained thereby. The starting material does not have to be copper sulfate crystal, and may be a material in which copper is dissolved with acid 15 containing sulfuric acid, or copper sulfate crystals manufactured therefrom.

BACKGROUND ART

Although copper sulfate (Cu_2SO_4) is white-colored powder, Hydrous 20 copper sulfate ($Cu_2SO_4 \cdot 5H_2O$) is the general term thereof, and this is an azurite blue crystal.

Copper sulfate is used as an electrolytic solution, pigment, insecticide, antiseptic, mordant, battery material, pharmaceutical and so on. In particular, when it is to be used as the electroplating solution in electronic components 25 such as a semiconductor device, high purity copper sulfate is being sought.

Commercially available copper sulfate has a purity level of 95 to 99.9wt%, and it is necessary to purify this even further to obtain a level of 4N to 5N or more.

As conventional technology, described is a method to obtain copper 30 sulfate with low Ni content by using electrolytic copper powder recovered

from an electrolytic solution via electrodeposition as the raw material, immersing this into an acid solution to selectively dissolve and remove Ni, dissolving the filtered copper powder in sulfuric acid, and subjecting this to crystallization (for example, c.f. Japanese Patent Laid-Open Publication No. 5 2001-10817).

Further, disclosed is technology for obtaining copper sulfate with low nickel content by employing an aqueous solution of copper sulfate containing nickel and heating this to 80°C or higher, collecting the copper sulfate crystal separated and sedimented, and concentrating this to effect recrystallization 10 (for example, c.f. Japanese Patent Laid-Open Publication No. 2001-31419).

DISCLOSURE OF THE INVENTION

An object of the present invention is to provide a manufacturing 15 method of high purity copper sulfate which allows the efficient removal of impurities from commercially available copper sulfate crystals at a low cost through dissolution with purified water and thermal concentration, and the high purity copper sulfate obtained thereby.

The present invention provides:

- 20 1. A manufacturing method of high purity copper sulfate, including the steps of dissolving copper sulfate crystals in purified water, performing evaporative concentration thereto, removing the crystals precipitated initially, performing further evaporative concentration to effect crystallization, subjecting this to filtration to obtain high purity copper sulfate, and 25 performing desiccation thereto;
2. A manufacturing method of high purity copper sulfate according to paragraph 1 above, wherein the initial pH of the solution in which the copper sulfate was dissolved in purified water is 2 to 4, and the pH of the solution after removing the crystals precipitated initially is 2 or less;
- 30 3. A manufacturing method of high purity copper sulfate according to

paragraph 1 or paragraph 2 above, wherein 10wt% or more of the initial crystal is removed in relation to the initial input;

4. A manufacturing method of high purity copper sulfate according to any one of paragraphs 1 to 3 above, wherein the filtration solution after the final

5 filtration is 2 to 40% of the original fluid volume;

5. A manufacturing method of high purity copper sulfate according to any one of paragraphs 1 to 4 above, wherein the desiccation temperature is 40 to 100°C;

6. High purity copper sulfate having a purity of 99.99wt% or higher, and in

10 which the content of transition metals such as Fe, Cr, Ni is 3wtppm or less;

7. High purity copper sulfate manufactured with the method according to any of paragraphs 1 to 5 above, wherein a purity is 99.99wt% or higher and in which the content of transition metals such as Fe, Cr, Ni is 3wtppm or;

8. High purity copper sulfate having a purity of 99.99wt% or higher, and in

15 which the content of Ag, Cl is respectively 1wtppm or less;

9. High purity copper sulfate according to paragraph 7 above, wherein the purity thereof is 99.99wt% or higher, and in which the content of Ag, Cl is respectively 1wtppm or less;

10. High purity copper sulfate having a purity of 99.99wt% or higher, and in

20 which the content of alkali metals such as Na, K and alkaline earth metals such as Ca, Mg is respectively 1wtppm or less;

11. High purity copper sulfate according to paragraph 9 above, wherein the purity thereof is 99.99wt% or higher, and in which the content of alkali metals such as Na, K and alkaline earth metals such as Ca, Mg is respectively

25 1wtppm or less;

12. High purity copper sulfate having a purity of 99.99wt% or higher, and in which the content of Si containing oxide is 10wtppm or less based on Si conversion; and

13. High purity copper sulfate according to paragraph 11 above, wherein

30 the purity thereof is 99.99wt% or higher, and in which the content of Si

containing oxide is 10wtppm or less based on Si conversion.

BRIEF DESCRIPTION OF THE DRAWINGS

5 The Figure is a diagram showing the flow of the manufacturing method
of high purity copper sulfate.

BEST MODE FOR CARRYING OUT THE INVENTION

10 The analytical values of representative impurities of commercially
available copper sulfate are shown in Table 1. Incidentally, by way of
reference, the comprehensive analytical values including other impurities are
shown in subsequent Table 6.

15 As shown in Table 1, this copper sulfate contains transition metals
such as iron, nickel, cobalt; and impurities such as Ca, Cr, Al, respectively, in
a content of roughly 0.5 to 100wtppm. Further, in addition to the above,
several wtppm to several ten wtppm of alkali elements such as Na, K; several
ten wtppb of impurities of radioactive elements such as U, Th; and foreign
matters such as oxides are also contained therein.

20 This copper sulfate is dissolved with purified water at room
temperature. When impurities such as organic matter or foreign matter are
mixed therein, impurities such as organic matter or foreign matter are subject
to filtration and removed by performing active carbon treatment. Thereupon,
noble metals such as Ag can also be removed.

25 When impurities such as organic matter or foreign matter are not mixed
therein, this active carbon treatment does not have to be performed. When
undissolved foreign matter exists, this may be removed with a filter cloth or
the like, but may also be removed upon removing the initial crystals in the
subsequent process. The initial pH of the solution in which the copper sulfate
30 was dissolved in purified water is made to be 2 to 4. The reason the pH of

this solution is made to be 2 to 4 is to promote the precipitation of initial crystals by ultrafine hydroxides of the copper sulfate being precipitated.

Next, the solution which dissolved copper sulfate or the filtration solution which filtrated copper sulfate solution is heated to a temperature of roughly 50 to 100°C, and subject to evaporative concentration.

TABLE 1

	Fe	Ni	Co	Ca	Cr	Al	Na	K	U	Th	wtppm
Raw Material	81	4.3	3.5	3.5	18.1	3.2	13	3.1	0.01	0.02	

After evaporative concentration, this is cooled to room temperature, and crystals are precipitated. These initial crystals contain numerous impurities, and, although the precipitated crystals are originally of a blue color, since the pH is high at 2 to 4, such crystals present a green color since they contain hydroxide. Further, numerous impurities can be removed as a result thereof. Thereupon, residue of the undissolved impurities can be simultaneously removed.

Moreover, by increasing the level of the foregoing evaporative concentration, the solution volume will decrease, the amount of initial crystals will increase, and the amount of removed impurities will increase. Nevertheless, when exceeding a certain level, the amount of removed impurities will not increase any further.

Thus, it is desirable to delete 10wt% or more of the initial crystals (low-grade copper sulfate containing impurities) in relation to the initial input. Further, the lower the purity of the raw material to be used, it is desirable that the amount of initial crystal to be removed is increased.

After filtering and removing the crystals precipitated initially, the pH of

the filtration solution is made to be 2 or less. This is to promote the precipitation of the copper sulfate without precipitating hydroxide. And, this is reheated to a temperature of 50 to 100°C, and subject to evaporative concentration.

5 Thereafter, this is cooled to room temperature, crystals of the copper sulfate are precipitated, and subject to filtration to obtain high purity copper sulfate of a blue color.

It is desirable that the filtration solution after the final filtration is 2 to 40wt% of the original fluid volume. In other words, the filtration solution is to
10 be the residual liquid containing non-crystallized copper sulfate, and the mixing of Na, K and so on which could not be removed from the initial crystal into the crystal can be prevented thereby. The desiccation temperature is preferably 40 to 100°C. If the temperature is less than 40°C, this is not preferable since much time will be required for removing the adsorption
15 moisture, and, if the temperature exceeds 100°C, this is not preferable since the Hydrous copper sulfate adhesive moisture will be removed and the form of the copper sulfate will change.

As a result of the foregoing process, the various impurities shown in Table 1 are reduced to 1wtppm or 0.1wtppm or less, respectively, and high
20 purity copper sulfate having a level of 4N to 5N or more can be obtained.

It is desirable that the content of Ag, Cl in the high purity copper sulfate is 1wtppm or less, respectively. This is because when copper sulfate is to be used as the plating solution, Ag, Cl will have an adverse effect on the plating, and will be concentrated in the film.

25 It is desirable that the content of alkali metals such as Na, K and alkaline earth metals such as Ca, Mg in the high purity copper sulfate is 1wtppm or less, respectively. This is because when copper sulfate is to be used as the plating solution in the manufacturing method of a semiconductor device, it is easily engulfed in the plating coating, and will have an adverse
30 on the performance of the semiconductor.

Further, it is desirable that the amount of Si (containing oxide and based on Si conversion) in the high purity copper sulfate is 10wtppm or less. This is because it will become foreign matter in the plating coating.

The flow of the manufacturing method of high purity copper sulfate
5 according to the present invention is shown in the Figure.

Examples

Examples of the present invention are now explained. Incidentally, these Examples are merely illustrative, and the present invention shall in no
10 way be limited thereby. In other words, the present invention shall include, within the scope of its technical spirit, any and all modes or modifications other than the Examples of this invention.

250g of commercially available copper sulfate ($\text{Cu}_2\text{SO}_4 \cdot 5\text{H}_2\text{O}$) having a [purity] level of 99.9wt% containing the impurities shown in Table 1 was
15 dissolved in 1000ml of purified water at room temperature.

Next, this was heated at 90°C to evaporate water in a prescribed amount. Next, this was cooled to room temperature, the initial crystals were precipitated, and subject to filtration.

The relationship of the amount of filtration solution and the weight of
20 initial crystals, as well as the relationship with the content of impurities (representatively showing the analytical value of Fe) contained in the ultimately refined material (copper sulfate) are shown in Table 2.

The weight of the initial crystals upon evaporating the filtration solution (liquid solution) such that the amount thereof decreases from 900ml to 600ml
25 increased from 1g to 80g together with such evaporation. Further, the content of Fe impurities decreased from 15ppm to 0.6ppm.

Nevertheless, the impurities did not decrease when removing the initial crystals exceeding 10wt%, and, since the yield of the copper sulfate will decrease, it is desirable to remove 10wt% or more of the initial crystal in
30 relation to the initial input.

TABLE 2

Solution Volume (ml)	Weight of Initial Crystal (g)	Content of impurities (Fe) ppm
900	1	15
800	25	3
700	50	0.8
600	80	0.6

Next, the filtration solution of the copper sulfate which improved purity
 5 obtained by removing the initial 50g of the crystal in Table 2 was reheated at
 90°C, and water was evaporated in a prescribed amount. Then, this was
 cooled to room temperature, the refined copper sulfate crystal was
 precipitated, and subject to filtration.

When the evaporation is increased drastically, there is a possibility that
 10 impurities such as Na and K will be mixed in the refined copper sulfate, and,
 therefore, evaporation was discontinued midway to obtain residual liquid.

The amount of filtration solution (solution volume), weight of the refined
 15 copper sulfate ($\text{Cu}_2\text{SO}_4 \cdot 5\text{H}_2\text{O}$) crystal, and the impurities (representatively
 shown with the Na content) contained in the refined copper sulfate crystal are
 shown in Table 3. As shown in Table 3, the solution volume after evaporation
 decreased from 300 to 10ml, and the refined copper sulfate crystal increased
 from 150 to 230g.

Further, the Na content in the refined copper sulfate crystal increased
 20 from 0.3 to 1.0ppm, and, when the refined copper sulfate crystal was 230g, it
 increased considerably to 5wtppm. Thus, excessive evaporation is not
 preferable from the perspective of impurities such as Na, K. Therefore, it is
 desirable that the filtration solution after the final filtration is 2 to 40% of the
 original fluid volume.

TABLE 3

Solution Volume (ml)	Weight of Refined Crystal (g)	Impurities contained in the Refined Crystal (Na) ppm
300	150	0.3
200	180	0.3
150	200	0.4
100	210	1.0
10	230	5

The analytical values of the representative impurities of the refined copper sulfate obtained in the refining process of the foregoing copper 5 sulfate when 300ml was evaporated in the process of removing the initial crystal, and this was concentrated to 300ml in the process of precipitating the refined copper sulfate crystal, are shown in Table 4. Incidentally, by way of reference, the comprehensive analytical values containing other impurities of the refined copper sulfate are shown in subsequent Table 7.

10 As shown in Table 4, the primary impurities are Fe: 0.8wtppm; Ni: 0.2wtppm; Co < 0.1wtppm; Ca < 0.1wtppm; Cr: 0.1wtppm; Al: 0.1wtppm; Na: 0.4wtppm; K < 0.1wtppm; U < 0.005wtppm; Th < 0.005wtppm and Si: 1.5wtppm, and, as evident upon comparison with Table 1, significant improvement in the purity has been confirmed with the simple evaporation 15 and filtration processes illustrated in the examples of the present invention.

These impurities are particularly disfavored in the copper electroplating to be performed to the likes of a circuit or wiring upon manufacturing a semiconductor device and so on, and the reduction of these impurities is extremely effective.

20 The initial make-up of the electroplating and electroless plating bath was conducted with this copper sulfate crystal, electrolytic copper plating and electroless copper plating were performed on a semiconductor wafer, and foreign matter (particles) and the embedding characteristics in the trench

were measured. Incidentally, a trench for embedding a copper wiring was formed on the surface of the semiconductor wafer, and a barrier metal formed from TaN was attached to the surface thereof. And a thin copper layer was attached thereon with sputtering or CVD. Plating was performed thereto.

5 The results are shown in Table 5.

Incidentally, the composition and plating conditions of the electroplating and electroless plating bath are as follows.

(Electroplating Bath and Plating Conditions)

10 As the plating solution, employed was copper sulfate: 20g/L (Cu); sulfuric acid: 200g/L; hydrochloric ion: 60mg/L (Cl); and additive [brightening agent, surface active agent] (Product Name: CC-1220, manufactured by Nikko Metal Plating): 1mL/L.

Plating conditions were plating bath temperature: 30°C; cathode current density: 2.0A/dm²; anode current density: 2.0A/dm²; and plating time:

15 1 (min).

(Electroless Plating Bath and Plating Conditions)

As the electroless plating solution, employed was copper sulfate: 4g/L (Cu); reducing agent: formaldehyde (37%) 3mL/L; complexing agent: EDTA - 2Na 30g/L; additive 1: Bipyridyl 20mg/L; and additive 2: polyethylene glycol 20mg/L.

20 And, plating was performed for 30 minutes at a plating temperature of 70°C and pH of 12.2.

TABLE 4

4-1

wtppm

	Fe	Ni	Co	Ca	Cr	Al
Refined Material	0.8	0.2	<0.1	<0.1	0.1	0.1

4-2

wtpm

	Na	K	U	Th	Si
Refined Material	0.3	<0.1	<0.005	<0.005	1.5

TABLE 5

	Plating Method	Amount of Foreign Matter (quantity)	Embedding Characteristics
Examples	Electroplating	3	Favorable
	Electroless Plating	1	Favorable
Comparative Examples	Electroplating	10	Inferior
	Electroless Plating	8	Inferior

5 Comparative Examples

Commercially available copper sulfate having a level of 99.9wt% containing the impurities shown in Table 1 was used to perform electroplating and electroless plating under the same conditions as the examples.

Foreign matter (particles) and the embedding characteristics in the trench in this case were measured as with Example 1, and the results are similarly shown in Table 5. As shown in Table 5, the amount of foreign matter in both electroplating and electroless plating in the comparative examples increased, and the embedding characteristics were inferior.

As shown in Table 5, in comparison to the comparative examples, the amount of foreign matter (particles) in the examples of the present invention was extremely small, there were no voids or engulfment of foreign matter, and a plating coating superior in embedding characteristics was obtained.

Accordingly, it is evident that the present invention yields a superior effect in that impurities can be effectively removed with a low cost by a relatively simple method of dissolving commercially available copper sulfate having a purity level of 95 to 99.9wt% in purified water and subjecting this to

thermal concentration. And, it has been confirmed thereby that transition metal elements, alkali metal elements and radioactive elements that are often contained in the copper sulfate can be reduced, and the high purity copper sulfate can be obtained efficiently.

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Effect of the Invention

The present invention yields a superior effect in that high purity copper sulfate can be manufactured at a low cost as a result of effectively removing impurities by dissolving commercially available copper sulfate having a purity

10 level of 95 to 99.9wt% in purified water and subjecting this to thermal concentration.

15

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25

30

TABLE 6
RAW MATERIAL OF COPPER SULFATE

ELEMENT	DENSITY (ppm wt)	ELEMENT	DENSITY (ppm wt)
Li	<0.01	Pd	<0.05
Be	<0.01	Ag	<0.05
B	1	Cd	<0.1
C	-	In	Binder
N	-	Sn	<1
O	Matrix	Sb	<0.5
F	<0.1	Te	<0.1
Na	1.3	I	<0.1
Mg	0.08	Cs	<0.1
Al	0.35	Ba	<0.05
Si	17	La	<0.05
P	2.3	Ce	<0.05
S	Matrix	Pr	<0.05
Cl	8.5	Nd	<0.05
K	<0.1	Sm	<0.05
Ca	<0.05	Eu	<0.05
Sc	<0.01	Gd	<0.05
Ti	0.04	Tb	<0.05
V	<0.01	Dy	<0.05
Cr	1.8	Ho	<0.05
Mn	0.02	Er	<0.05
Fe	21	Tm	<0.05
Co	0.01	Yb	<0.05
Ni	0.08	Lu	<0.05
Cu	Matrix	Hf	<0.05
Zn	<0.1	Ta	<5
Ga	<0.01	W	<0.05
Ge	<0.05	Re	<0.01
As	<0.05	Os	<0.01
Se	<0.1	Ir	1.8
Br	<0.1	Pt	<0.05
Rb	<0.01	Au	<0.1
Sr	<0.01	Hg	<0.05
Y	<0.01	Tl	<0.01
Zr	<0.05	Pb	=<1.5
Nb	<0.05	Bi	<0.05
Mo	<0.05	Th	<0.005
Ru	<0.05	U	<0.005
Rh	<100		

TABLE 7
HIGH PURITY COPPER SULFATE

ELEMENT	DENSITY (ppm wt)	ELEMENT	DENSITY (ppm wt)
Li	<0.01	Pd	<0.05
Be	<0.01	Ag	<0.05
B	0.04	Cd	<0.1
C	-	In	Binder
N	-	Sn	<1
O	Matrix	Sb	<0.5
F	<0.1	Te	<0.1
Na	0.4	I	<0.1
Mg	<0.01	Cs	<0.1
Al	0.1	Ba	<0.05
Si	1.5	La	<0.05
P	0.1	Ce	<0.05
S	Matrix	Pr	<0.05
Cl	0.8	Nd	<0.05
K	<0.1	Sm	<0.05
Ca	<0.1	Eu	<0.05
Sc	<0.01	Gd	<0.05
Ti	<0.01	Tb	<0.05
V	<0.01	Dy	<0.05
Cr	0.1	Ho	<0.05
Mn	0.03	Er	<0.05
Fe	0.8	Tm	<0.05
Co	<0.01	Yb	<0.05
Ni	0.2	Lu	<0.05
Cu	Matrix	Hf	<0.05
Zn	0.1	Ta	<5
Ga	<0.01	W	<0.05
Ge	<0.05	Re	<0.01
As	<0.05	Os	<0.01
Se	<0.1	Ir	<0.05
Br	<0.1	Pt	<0.05
Rb	<0.01	Au	<0.1
Sr	<0.01	Hg	<0.05
Y	<0.01	Tl	<0.01
Zr	<0.05	Pb	=<0.7
Nb	<0.05	Bi	<0.05
Mo	<0.05	Th	<0.005
Ru	<0.05	U	<0.005
Rh	<100		